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11. Describe the sublevels and orbitals that constitute the third and fourth principal energy levels of hydrogen. What are the general shapes of their probability maps? 12. How does electron spin affect the total number of electrons that can be accommodated in a given orbital? 13.

### Chapter 11: Basic Review Worksheet

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Sample Study Sheet 11.2: Abbreviated Electron Configurations To get a review of the most important topics in the chapter, fill in the blanks in the Key Ideas section. Work all of the selected problems at the end of the chapter, and check your answers with the solutions provided in this chapter of the study guide. Ask for help if you need it.

### Chapter 11 Modern Atomic Theory - An Introduction to Chemistry

This means that 100.00 g of sucrose always contains 42.11 g of carbon, 6.48 g of hydrogen, and 51.41 g of oxygen. First we will use the molecular formula of sucrose (C<sub>12</sub>H<sub>22</sub>O<sub>11</sub>) to calculate the mass percentage of the component elements; then we will show how mass percentages can be used to determine an empirical formula.

### Chapter 11.2: Empirical and Molecular Formulas - Chemistry ...

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