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The empirical formula of a compound represents the simplest whole-number ratio between the elements that make up the compound. This 10-question practice test deals with finding empirical formulas of chemical compounds. A periodic table will be required to complete this practice test. Answers for the test appear after the final question:

Empirical Formula Practice Test Questions

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Empirical Formula Study Guide With Answer Sheet

SHOW WORK ON A SEPARATE SHEET OF PAPER. Write the empirical formula for the following compounds. 1) C_6H_6 . 2) C_8H_{18} . 3) WO_2 . 4) $C_2H_6O_2$. 5) X_3Y_{13} . 6) A compound with an empirical formula of C_2OH_4 and a molar mass of 88 grams per mole.

Empirical and Molecular Formula Worksheet

Empirical Formula Examples. Glucose has a molecular formula of $C_6H_{12}O_6$. It contains 2 moles of hydrogen for every mole of carbon and oxygen. The empirical formula for glucose is CH_2O . The molecular formula of ribose is $C_5H_{10}O_5$, which can be reduced to the empirical formula CH_2O .

Empirical Formula: Definition and Examples

Work sheet on Empirical formula answer sheet given with extension material

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Empirical Formula 1) What is the empirical formula of a compound that contains 0.783g of Carbon, 0.196g of Hydrogen and 0.52 lg of Oxygen? 2) What is empirical formula of a compound which consists of 89.14% Au and 10.80% of O? 3) What is empirical formula if compound consists of 21.20/oN, 6.10/01-1, 24.2%S and 48.50/00? Molecular Formula

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D Determine the molecular formula for each compound described. 9. A compound has an empirical formula of NO_2 and a molar mass of 92.02 g/mol . 10. A compound has an empirical formula of $\text{C}_2\text{H}_3\text{O}$ and a molar mass of 172 g/mol . 11. Ibuprofen, a common headache remedy, has an empirical formula of $\text{C}_{13}\text{H}_{18}\text{O}_2$ and a molar mass of approximately 215 g/mol . 12.

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Empirical and Molecular Formula Worksheet . SHOW WORK ON A SEPARATE SHEET OF PAPER. Write the empirical formula for the following compounds. 1) C_6H_6 . 2) C_8H_{18} . 3) WO_2 . 4) $\text{C}_2\text{H}_6\text{O}_2$. 5) $\text{X}_{39}\text{Y}_{13}$.

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