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6.2 6.2.1 Vinegar is a dilute form of acetic acid. A sample of acetic acid has the following percentage composition: 39,9% carbon 6,7% hydrogen 53,4% oxygen Determine the molecular formula of acetic acid if the molar mass of acetic acid is $60 \text{ g} \cdot \text{mol}^{-1}$. (6) 6.2.2

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Give a reason why acetic acid is regarded as a monoprotic acid. (2)

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(b) The relative atomic mass (A_r) of chlorine is given on the Periodic Table as 35.5. Explain why the relative atomic mass is not a whole number.

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