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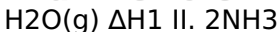
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ones of the following  
reactions are  
endothermic in other

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words  $\Delta H$  is positive? I.



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change. Gibbs free energy and spontaneity. Gibbs free energy example. More rigorous Gibbs free energy / spontaneity relationship.

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temperature of  
20°C. The specific  
heat of A is larger than  
that of B.

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ANSWER KEY 1.

Calculate the standard  
enthalpy of reaction for  
the

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6) 1. Consider 2 metals, A and B, each having a mass of 100 g and an initial temperature of 20 °C. The specific heat of A is larger than that of B. Under the same heating conditions, which metal would take longer to reach 21 °C? Explain your reasoning. 2.

**Thermo PRACTICE  
PROBLEMS**

Answer: The n

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number ( )ordinati01  
wing data: equilibrium  
.ch of sOC. If? Br<sub>2</sub>, the  
cc xtively an 2560C ing  
the follc 2(g) will be  
5620C very slow. .2  
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AP Chemistry Practice  
Test, Ch. 6:  
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Thermochemistry  
Name \_\_\_\_\_ MULTIPLE  
CHOICE. Choose the  
one alternative that  
best completes the  
statement or answers  
the question. 1) A  
chemical reaction that  
absorbs heat from the  
surroundings is said to  
be \_\_\_\_\_ and has a  
\_\_\_\_\_  $\Delta H$  at constant  
pressure.

A) endothermic,  
positive

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**Practice Test, Ch. 6:  
Thermochemistry ...**

Answers,  
Thermochemistry  
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6. When 26.7 g of H<sub>2</sub>S  
was burned in excess  
oxygen, 406 kJ was  
released. What is H for  
the following

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thermodynamics. First  
law of thermodynamics  
problem solving. PV  
diagrams - part 1:  
Work and isobaric  
processes. PV



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diagrams - part 2:

Isothermal, isometric,  
adiabatic processes.

Second law of  
thermodynamics.

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Thermochemistry  
practice problems 1)

How can energy be  
transferred to or from a  
system? A) Energy can  
only be transferred as  
potential energy being  
converted to kinetic

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energy. B) Energy can be transferred only as heat. C) Energy can be transferred only as work. D) Energy can be transferred as heat and/or work.

## **Chemistry @ POB - Home**

Thermochemistry  
Practice Problems -  
Answers 1. What will be  
sign for  $q$  and  $W$  if an  
isolated system absorb  
energy from the  
surrounding and does

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work for expansion. 2.

The amount of work done in joules by the system in expanding from 1.50L to 2.3L against a constant atmospheric pressure of about 1.3atm. 3.

**1. 2 3. -**

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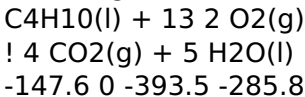
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calculate the quantity  
of heat transferred u...

## **Calorimetry Problems, Thermochemistry Practice, Specific ...**

Thermochemistry Problems, Page 6 of 7, [9]



1)  $\Delta H$

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Chapter Exam

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your answers to the

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questions and click  
'Next' to see the next  
set of questions.

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For each of the  
following questions or  
statements, select the  
most appropriate  
response and click its  
letter:

## **Quiz #3-3 PRACTICE: Thermochemistry |**



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to the concepts in your  
lecture. These are just  
problems for extra  
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